

TRANSMITTAL LETTER TO THE UNITED STATES
DESIGNATED/ELECTED OFFICE (DO/EO/US)
CONCERNING A FILING UNDER 35 U.S.C. 371

G80-019 US

U.S. APPLICATION NO. (if known) 37 CFR 1.1

09/744655

INTERNATIONAL APPLICATION NO.

PCT/ES99/00242

INTERNATIONAL FILING DATE

July 29, 1999

PRIORITY DATE CLAIMED

July 31, 1998

TITLE OF INVENTION

PRODUCT FOR DESACIDIFICATION OF CELLULOSE MATERIAL, PRODUCTION
AND UTILIZATION THEREOF

APPLICANT(S) FOR DO/EO/US

Area 1 GUERRA

Applicant herewith submits to the United States Designated/Elected Office (DO/EO/US) the following items and other information:

1. ☒ This is a FIRST submission of items concerning a filing under 35 U.S.C. 371.
2. ☐ This is a SECOND or SUBSEQUENT submission of items concerning a filing under 35 U.S.C. 371.
3. ☒ This express request to begin national examination procedures (35 U.S.C. 371(f)) at any time rather than delay examination until the expiration of the applicable time limit set in 35 U.S.C. 371(b) and PCT Articles 22 and 39(1).
4. ☒ A proper Demand for International Preliminary Examination was made by the 19th month from the earliest claimed priority date.
5. ☒ A copy of the International Application as filed (35 U.S.C. 371(c)(2))
 - a. ☐ is transmitted herewith (required only if not transmitted by the International Bureau).
 - b. ☒ has been transmitted by the International Bureau.
 - c. ☐ is not required, as the application was filed in the United States Receiving Office (RO/US)
6. ☒ A translation of the International Application into English (35 U.S.C. 371(c)(2)).
7. ☒ Amendments to the claims of the International Application under PCT Article 19 (35 U.S.C. 371(c)(3))
 - a. ☐ are transmitted herewith (required only if not transmitted by the International Bureau).
 - b. ☐ have been transmitted by the International Bureau.
 - c. ☐ have not been made; however, the time limit for making such amendments has NOT expired.
 - d. ☒ have not been made and will not be made.
8. ☐ A translation of the amendments to the claims under PCT Article 19 (35 U.S.C. 371(c)(3)).
9. ☐ An oath or declaration of the inventor(s) (35 U.S.C. 371(c)(4)).
10. ☒ A translation of the annexes to the International Preliminary Examination Report under PCT Article 36 (35 U.S.C. 371(c)(5)).

Items 11. to 16. below concern other document(s) or information included:

11. ☐ An Information Disclosure Statement under 37 CFR 1.97 and 1.98.
12. ☐ An assignment document for recording. A separate cover sheet in compliance with 37 CFR 3.28 and 3.31 is included.
13. ☒ A FIRST preliminary amendment.
☐ A SECOND or SUBSEQUENT preliminary amendment.
14. ☐ A substitute specification.
15. ☐ A change of power of attorney and/or address letter.
16. ☐ Other items or information:

16(a) Please check appropriate box:

☐ CHAPTER I - 20 Mos.☒ CHAPTER II - 30 Mos.

VIA EXPRESS MAIL

LABEL NO.: EL769434601 US

POSTED 554460

U.S. APPLICATION NO. (U.S. PATENT & TRADEMARK OFFICE)

INTERNATIONAL APPLICATION NO.

ATTORNEY'S DOCKET NUMBER

09/744655

PCT/ES99/00242

G80-019US

17. ☒ The

Basic National Fee (37 CFR 1.492(a)(1)-(5)):

Search Report has been prepared by the EPO or JPO.....

International preliminary examination fee paid to USPTO (37 CFR 1.482)

No international preliminary examination fee paid to USPTO (37 CFR 1.482)
but international search fee paid to USPTO (37 CFR 1.445(a)(2))..Neither international preliminary examination fee (37 CFR 1.482) nor
international search fee (37 CFR 1.445(a)(2)) paid to USPTO.....International preliminary examination fee paid to USPTO (37 CFR 1.482) -
and all claims satisfied provisions of PCT Article 33(2)-(4).....

ENTER APPROPRIATE BASIC FEE AMOUNT = \$ 1,000.00

Surcharge of \$130.00 for furnishing the oath or declaration later than ☐ 20 ☐ 30
months from the earliest claimed priority date (37 CFR 1.492(e)).

Claims	Number Filed	Number Extra	Rate	
Total Claims	-20 =		X	\$
Independent Claims	-3 =		X	\$
Multiple dependent claims(s) (if applicable)			+	\$

TOTAL OF ABOVE CALCULATIONS = \$ 1,000.00

Reduction by 1/2 for filing by small entity, if applicable. Verified Small Entity
Status is hereby verified by the undersigned.

SUBTOTAL = \$ 500.00

Processing fee of \$130.00 for furnishing the English translation later the ☐ 20 ☐ 30
months from the earliest claimed priority date (37 CFR 1.492(f)).

TOTAL NATIONAL FEE = \$ 500.00

Fee for recording the enclosed assignment (37 CFR 1.21(h)). The assignment must be
accompanied by an appropriate cover sheet (37 CFR 3.28, 3.31). \$40.00 per property +

TOTAL FEES ENCLOSED = \$ 500.00

Amount to be:
refunded \$
charged \$

- a. ☒ A check in the amount of \$ 500.00 to cover the above fees is enclosed.
- b. ☐ Please charge my Deposit Account No. _____ in the amount of \$ _____ to cover the above fees.
A duplicate copy of this sheet is enclosed.
- c. ☒ The Commissioner is hereby authorized to charge any additional fees which may be required, or credit any
overpayment to Deposit Account No. 14-1431. A duplicate copy of this sheet is enclosed.

NOTE: Where an appropriate time limit under 37 CFR 1.494 or 1.495 has not been met, a petition to revive (37 CFR 1.137(a) or (b)) must be filed and granted to restore the application to pending status.

SEND ALL CORRESPONDENCE TO:

NOTARO & MICHALOS P.C.
100 Dutch Hill Road, Suite 110
Orangeburg, NY 10962-2100

SIGNATURE

Peter C. MICHALOS

NAME

09/744655

JC02 Rec'd PCT/PTO 29 JAN 2001

CERTIFICATE OF EXPRESS MAILING

I hereby certify that this correspondence is being deposited with the U.S. Postal Service as Express Mail Label No. EL769434601 US in an envelope addressed to Box Non-Fee Amendment, Commissioner for Patents, Washington, D.C. 20231 on January 29, 2001

By: Gemmae Berke

Dated: January 29, 2001

Patent

Atty. Docket: G80-019 US

IN THE UNITED STATES PATENT AND TRADEMARK OFFICE

Applicant : Areal Guerra
PCT Serial No. : PCT/ES99/00242
Filing Date : January 29, 2001
For : PRODUCT FOR DESACIDIFICATION OF
CELLULOSE MATERIAL, PRODUCTION AND
UTILIZATION THEREOF
Examiner : --
Group Art Unit : --

Commissioner for Patents
Washington, D.C. 20231

PRELIMINARY AMENDMENT

Sir:

Preliminary to the examination of the above-identified patent application, please amend the application as follows.

IN THE CLAIMS:

Please amend the claims, as follows:

Claim 5, line 3, change "any of claims 1 to 4" to --claim 1--.

Claim 9, line 5, change "any of the claims 1 to 4" to --claim 1--.

REMARKS:

Claims 1-10 are in the case and presented for consideration.

The application has been amended to place the same in better form for U.S. practice. Specifically, the multiple dependent claims have been amended to depend from a single claim.

Accordingly, the application and claims are believed to be in condition for allowance, and favorable action is respectfully requested.

Respectfully submitted,



Peter C. Michalos
Reg. No. 28,643
Attorney for Applicants
(914) 359-7700

Dated: January 29, 2001

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100 Dutch Hill Road, Suite 110
Orangeburg, New York 10962-2100

09/744655

Attorney's Docket No. G80-019 US

Applicant or Patentee R. Areal Guerra
 Serial or Patent No.: 09/744,655
 Filed or Issued: January 29, 2001
 For: PRODUCT OF DEACIDIFICATION...

VERIFIED STATEMENT (DECLARATION CLAIMING SMALL ENTITY STATUS)
 (37 CFR 1.9(f) and 1.27(c) - SMALL BUSINESS CONCERN)

I hereby declare that I am

☒ the owner of the small business concern identified below;
 an official of the small business concern empowered to act on behalf
 of the concern identified below:

NAME OF SMALL BUSINESS CONCERN UNIVERSITAT POLITÈCNICA DE CATALUNYA
 ADDRESS OF CONCERN COLON 11, 08222 TERRASSA, SPAIN

I hereby declare that the above-identified small business concern qualifies as a small business concern as defined in 13 CFR 121.12, and reproduced in 37 CFR 1.9(d), for purposes of paying reduced fees to the United States Patent and Trademark Office, in that the number of employees of the concern, including those of its affiliates, does not exceed 500 persons. For purposes of this statement, (1) the number of employees of the business concern is the average over the previous fiscal year of the concern of the persons employed on a full-time, part-time or temporary basis during each of the pay periods of the fiscal year, and (2) concerns are affiliates of each other when either, directly or indirectly, one concern controls or has the power to control the other, or a third party or parties controls or has the power to control both.

I hereby declare that rights under contract or law have been conveyed to and remain with the small business concern identified above with regard to the invention, entitled SAME AS ABOVE
 by inventor(s) ROGELIO AREAL GUERRA
 described in

☒ the specification filed herewith
 application serial no. 09/744,645, filed 1-29-01
 patent no. _____, issued _____

If the rights held by the above-identified small business concern are not exclusive, each individual, concern or organization having rights to the invention is listed below and no rights to the invention are held by any person, other than the inventor, who could not qualify as an independent inventor under 37 CFR 1.9(c), if that person made the invention, or by any concern which would not qualify as a small business concern under 37 CFR 1.9(d) or a nonprofit organization under 37 CFR 1.9(e). NOTE: Separate verified statements are required from each named person, concern or organization having rights to the invention averring to their status as small entities. (37 CFR 1.27)

NAME _____
 ADDRESS _____
☐ INDIVIDUAL ☐ SMALL BUSINESS CONCERN ☐ NONPROFIT ORGANIZATION

NAME _____
 ADDRESS _____
☐ INDIVIDUAL ☐ SMALL BUSINESS CONCERN ☐ NONPROFIT ORGANIZATION

I acknowledge the duty to file, in this application or patent, notification of any charge in status resulting in loss of entitlement to small entity status prior to paying, or at the time of paying, the earliest of the issue fee or any maintenance fee due after the date on which status as a small entity is no longer appropriate. (37 CFR 1.28(b)).

I hereby declare that all statements made herein of my own knowledge are true and that all statements made on information and belief are believed to be true; and further that these statements were made with the knowledge that willful false statements and the like so made are punishable by fine or imprisonment, or both, under section 1001 of Title 18 of the United States Code, and that such willful false statements may jeopardize the validity of the application, any patent issuing thereon, or any patent to which this verified statement is directed.

NAME OF PERSON SIGNING JAUME PAGÈS
 TITLE OF PERSON OTHER THAN OWNER RECTOR
 ADDRESS OF PERSON SIGNING CTJORD. GIRONA 31 08024 BARCELONA
 SIGNATURE [Signature] DATE 29/01/01

PRODUCT FOR THE DE-ACIDIFICATION OF CELLULOSE MATERIAL,
AND PRODUCTION AND USE THEREOF

FIELD OF THE INVENTION

5 This invention relates to the preservation of
objects that contain cellulose-type materials, in
particular, to a product, in the form of a solution, for
de-acidification of cellulose-type materials composed of
carbonated magnesium di-n-propylate, n-propanol and a
10 hydrofluorocarbon diluent.

BACKGROUND OF THE INVENTION

There is a lot of concern for the deterioration
undergone by cellulose-type materials (books, documents,
15 maps, press, magazines and tissues) held in libraries,
newspaper archives, museums and archives. This
deterioration endangers the physical integrity of many
documents, books and tissues and thus forces a
restriction in access and handling thereof. The most
20 significant cause of deterioration in cellulose-type
materials is hydrolysis catalyzed by acids in the
cellulose, giving rise to a shortening of the polymeric
chains of the cellulose, which, in turn, leads to a
weakening of the cellulose fibers and they become more
25 fragile. Many factors contribute to the acidification
of cellulose-type materials. For example, acidic gases
present in contaminated air, the use of alum and
colophony (abietic acid) in paper production, products
for oxidative degradation of cellulose, inks and other
30 substances that oxidize cellulose which are present in
paper and books.

A great deal of effort has been put in to the
development of processes to achieve de-acidification of
35 cellulose-type materials and thus to prolong their
useful life by stopping acidic hydrolysis degradation
processes. Specific compositions and products have been

studied for de-acidification, capable of neutralizing the acid present in cellulose-type materials and introducing an alkaline reserve, which inhibits re-acidification. Technology has moved on from expensive craft methods using aqueous medium to the development of methods for applying alkaline reagents, dissolved in non-aqueous vehicles, either in liquid or solid state, to prevent books from falling apart.

10 The North American patent US 3.676.182 (R.D. Smith, 11th July 1972) relates to a procedure for non-aqueous de-acidification of paper using magnesium methoxide, dissolved in methanol to a concentration lying between 5 % and 11 %, as the de-acidifying agent. The resulting solution can be diluted with a chlorofluorocarbon (CFC) diluent, such as trichlorofluoromethane or dichlorodifluoromethane, until a liquid is formed, under pressure, that contains between 1 % and 2 % of the de-acidifying agent. The CFC compound quickly evaporates from paper impregnated with the de-acidifying solution containing methanol, thus minimizing the solvent effect of methanol on certain inks. The books and papers can be treated by immersion in the de-acidifying solution, or the solution can be applied with a brush or using a spray. Nevertheless, this method is not free from drawbacks as the magnesium methoxide is extremely sensitive to water. So much so that even traces of moisture lead to immediate hydrolysis forming a gelatinous precipitate of magnesium hydroxide. This is insoluble in water and many organic solvents. Unless the paper has been sufficiently desiccated before treatment premature hydrolysis of the magnesium methoxide may occur, leading to the formation of unwanted white marks of powdered magnesium hydroxide on the surface of the paper, giving the paper a rough feel. On the other hand, the spray nozzles frequently become blocked and the brush bristles become caked, making it necessary to

clean them and then dry them thoroughly before using them again.

The North American Patent US 3.939.091 (Kelly, G.B., 7th February 1976) discloses and claims a product for the de-acidification of paper that overcomes the problems of previous de-acidification solutions based on magnesium methoxide dissolved in CFC. The magnesium methoxide dissolved in methanol or in a methanol and trichlorotrifluoroethane (Freon® F7) mixture is reacted with gaseous carbon dioxide to obtain a colourless 20 % solution of methoxymagnesium methylcarbonate. The resulting solution is more tolerant of water than the magnesium methoxide solutions, and obstructions are not produced in the spray nozzles as often as before.

The Canadian patent CA 1.147.510 (Smith, R.D., 7th of June 1983) discloses a method for the production of methoxymagnesium methylcarbonate. Metallic magnesium is allowed to completely react with methanol to form magnesium methoxide. This is then re-dissolved in methanol saturated with carbon dioxide forming a methoxymagnesium methylcarbonate solution. This solution is diluted with trichlorotrifluoro-ethane or with dichlorodifluoromethane. The diluted solution is sprayed onto the paper to be de-acidified. As an alternative, the paper can be submerged in the same solution, with the liquid state being maintained under pressure.

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The North American patent US 4.860.685 (Smith, R.D., 29th of August 1989) and the corresponding Canadian patent CA 1.272.018 (31st of July 1990) describe a flexible spraying system for the de-acidification of cellulose-type materials. A de-acidifying agent, preferable an alkoxide of magnesium carbonate, is combined with a CFC diluent, preferably

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trichlorotrifluoroethane, and with a propellant gas, preferably dichlorodifluoromethane. Additional pressurization and propulsion can be achieved using an inert gas such as nitrogen. The carbonated magnesium alkoxide, such as methoxymagnesium methylcarbonate or ethoxymagnesium ethylcarbonate, is produced by dissolving the corresponding magnesium alkoxide in alcohol in the presence of carbon dioxide.

The procedures described above use CFCs as these are essentially non-reactive and so do not imply a direct toxic risk for living beings. However, these same properties of chemical inertness that the CFCs possess make them problematic in the stratosphere. It has been demonstrated that the CFCs are the main factor responsible for the destruction of the stratospheric ozone layer and that they contribute to the greenhouse effect in the troposphere. The problem revolves around the chlorine that is liberated when the CFCs are exposed to ultraviolet radiation and other more penetrating radiations in the stratosphere. The chlorine destroys ozone by catalysing its transformation into molecular oxygen. The production of CFCs has been banned since the 1st of January 1996 (Copenhagen Agreement which amends the Montreal Protocol of the 23-25th of November 1992), and use thereof should be totally eliminated before the year 2030.

The Canadian patent CA 2.142.195 (Worsford, D. James, 8th of February 1995) discloses and claims a product for de-acidification of cellulose-type materials which consists of a de-acidifying reagent, such as methoxymagnesium methylcarbonate or ethoxymagnesium ethylcarbonate, a solvent such as methanol or ethanol, and a hydrochlorofluorocarbon diluent (HCFC) or hydrofluorocarbon (HFC). The product patented can be applied to the cellulose-type material by spray or with

a brush, or by immersing the cellulose-type material in the product. As a novelty with respect to the patents cited earlier this patent describes the use of diluents that are not prohibited by the Montreal Protocol [HCFCs and HFCs].

The use of HCFCs and HFCs as a substitute for CFCs substantially reduces the number of chlorine radicals that can be liberated in the stratosphere on exposure to cosmic radiation, thus reducing the potential for ozone destruction. On the other hand, the presence of hydrogen means that the compound is more reactive, such that it tends to decompose more easily before reaching the stratosphere. The rate of decomposition of HCFCs and HFCs is higher than that of the CFCs, whose lifetime is quoted at being between 60 and 100 years.

The object of the present invention is to find a new product for the non-aqueous de-acidification of cellulose-type materials, thus providing an addition to the arsenal of means available for combating acidification of cellulose-type materials.

Therefore, an object of this invention consists of a product, in the form of a solution, for the non-aqueous de-acidification of cellulose-type materials which comprises carbonated magnesium di-n-propylate, n-propanol and an HFC diluent.

An additional object of this invention consists of a procedure for production of said product for the de-acidification of cellulose-type material.

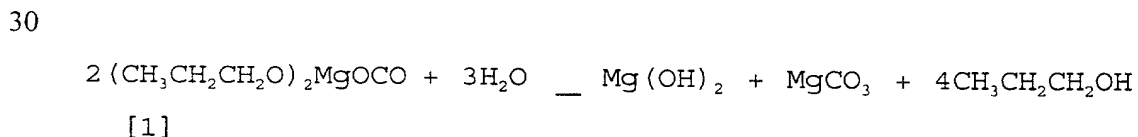
Another additional object of this invention consists of a method for de-acidification of cellulose-type material which comprises use of the aforementioned product.

DETAILED DESCRIPTION OF THE INVENTION

The present invention provides a product for de-acidification of cellulose-type material, hereinafter
 5 denominated the product of the invention, characterized because it is presented in the form of a solution and is comprised of:

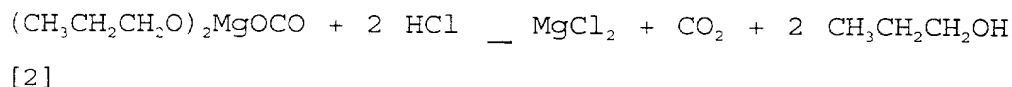
- carbonated magnesium di-n-propylate,
- n-propanol, and
- 10 - an HFC diluent selected from 1,1,1,2-tetrafluoroethane (HFC 134a) y 1,1,1,2,3,3,3-heptafluoropropane (HFC 227).

Carbonated magnesium di-n-propylate is a white solid, with the formula $(\text{CH}_3\text{CH}_2\text{CH}_2\text{O})_2\text{MgOCO}$, soluble in
 15 anhydrous n-propanol and in medium-polarity solvents, such as HFC 227 and HFC 134a, whose polarities are 0.8 Debyes (D). The meaning of the term "organic solvents of medium polarity" in this description includes solvents with a dipole moment (μ) less than water [:
 20 1.8 D], for example, toluene [: 0.40 D] and n-propanol [: 1.5 D]. carbonated magnesium di-n-propylate has a dipole moment of 0.8 D and so, in general, it is soluble in any proportion in the HFC diluent used for producing the product of the invention. The dipole moment of this
 25 diluent is approximately 0.9 D. Carbonated magnesium di-n-propylate, on coming into contact with moist air, reacts to give magnesium carbonate and magnesium hydroxide, and slowly releases n-propanol according to the reaction [1]:



Carbonated magnesium di-n-propylate reacts
 35 vigorously with mineral acids releasing carbon dioxide and n-propanol, forming soluble magnesium salts as

indicated by reaction [2]:



5

As a result of the high solubility of carbonated magnesium di-n-propylate in polar solvents, characterized by low surface tension and viscosity values, it can be dispersed deeply through the pores in the paper, reaching all points of cellulose-type material, and neutralizing the acids present. The excess reagent slowly decomposes into magnesium carbonate hydroxide, a stable compound of the formula $(\text{MgCO}_3)_4.\text{Mg}(\text{OH})_2.5\text{H}_2\text{O}$, which it deposited among the paper fibres to act as a alkaline reservoir protecting against acidic atmospheric contaminants and acids released in the paper during ageing. The trials carried out on papers treated and submitted to accelerated ageing have shown that treatment with the product of the invention [see Example 3] stabilizes the properties of the mechanical resistance of the paper, manifesting itself as a stabilization in the degree of whiteness.

Carbonated magnesium di-n-propylate can be obtained from magnesium di-n-propylate by a procedure that comprises the stages of:

- reacting a suspension of magnesium di-n-propylate in n-propanol with gaseous anhydrous carbon dioxide, until a solution of carbonated magnesium di-n-propylate is obtained; and

- separating said solution of carbonated magnesium di-n-propylate in n-propanol from the residual products, for example by decantation.

Reaction of magnesium di-n-propylate with anhydrous carbon dioxide is an exothermic reaction, reaching

temperatures of up to 45° C and the initial solid compound insoluble in n-propanol [magnesium di-n-propylate] is transformed into a soluble form consisting of carbonated magnesium di-n-propylate. The reaction is
 5 taken to have reached its conclusion when the temperature drops to room temperature. Next, the dark solid insoluble particles are allowed to sediment out on the bottom of the recipient and to give a clear and transparent solution. The resulting solution is
 10 collected by conventional techniques, for example, by decantation or, preferably, by suction under vacuum, and transferred to recipients suitable for loading, dilution or dosing. The carbonated magnesium di-n-propylate obtained is purified, for example, by evaporation of
 15 part of the solvent (typically 20-30 %) under vacuum.

The concentration of carbonated magnesium di-n-propylate in the alcoholic solution lies between 30 and 70 % (P/P), preferably between 45 and 50 % (P/P) of
 20 carbonated magnesium di-n-propylate, the rest consisting of n-propanol. It can be adjusted to the desired concentration by conventional methods, for example, by dilution with n-propanol or by elimination of excess solvent.

25

Magnesium alkoxides can be obtained using known methods, for example, *Metal Alkoxides*, by Bradley, D.C., Mehrotra, R.C. and Gaur, D.P., Academic Press, London (1978), and the work of Thoms, H., Eppele, M., Viebrock,
 30 H. and Reller, A., *J. Mater. Chem.* 5(4)589, (1995), where the synthesis of different magnesium alkoxides from alcohols of up to four carbon atoms is described. Although various magnesium alkoxides are known, the literature only describes carbonated magnesium alkoxides
 35 derived from methanol (methoxymagnesium methylcarbonate) and ethanol (ethoxymagnesium ethylcarbonate) and the carbonate of methoxy and butoxy polyethylene glycol [WO

90/03466].

The suspension of magnesium di-n-propylate in n-propanol can be prepared by different procedures. One of them (Procedure A) comprises reacting magnesium metal with anhydrous n-propanol, in the presence of iodine, at the temperature corresponding to the boiling point of the mixture. Another procedure (Procedure B) comprises reacting magnesium metal with anhydrous n-propanol in the presence of iodine at reflux temperature and the addition of toluene to form an azeotrope with n-propanol. The magnesium metal used in either Procedure A or B may in the form of a strip, in which case, it requires a suitable preparation (see Example 1.2).

Alternatively, the suspension of magnesium di-n-propylate in n-propanol can be prepared by a procedure (Procedure C) which does not require the use of a strip of magnesium but rather magnesium in powder form, with a granulometric distribution between 50 and 150 m. This procedure C therefore comprises of; (i) reacting magnesium in powder form, with a granulometric distribution between 50 and 150 m, with anhydrous n-propanol in the presence of iodine, which acts as a catalyst. The mixture is gently heated until hydrogen begins to be released. From this moment on, as the reaction is strongly exothermic, the next step (ii) consists of cooling the reaction mixture to the temperature corresponding to the boiling point, with a view to controlling the reaction rate with gentle boiling until hydrogen is no longer given off and the last remaining particles of magnesium have disappeared. In this fashion a suspension of magnesium di-n-propylate in n-propanol is obtained.

The use of magnesium in the form of powder with the indicated granulometric distribution (50-150 m) means

that the reaction of magnesium with n-propanol in the presence of iodine is exothermic, and so the reaction medium should be cooled instead of additional energy being provided. This also allows the reaction time to
5 be reduced [typically, the total reaction time for production of magnesium di-n-propylate is 4-5 hours].

Carbonated magnesium di-n-propylate can be used in the production of a product, in the form of a solution,
10 suitable for de-acidification of cellulose-type material along with n-propanol, and a diluent selected from HFC 134a and HFC 227.

The meaning of "cellulose-type material" as used
15 in this description refers to a material totally or partially composed of cellulose fibres, including paper of all classes, tissues and fabrics which contain cellulose fibres of vegetal origin, for example, those produced from wood, cotton, flax, jute, hemp and other
20 plants. Examples of such cellulose-type materials include books, documents, maps, works of art, articles elaborated with said materials, clothing, flags, etc.

For use as a reagent for de-acidification of
25 cellulose-type material a solution of carbonated magnesium di-n-propylate and n-propanol is produced, in HFC diluent, at a concentration lying between 30 % and 70 % (W/V) for application by machine and of 3.5 % to 4.5 % (W/V) in said HFC diluent for application by
30 sprays. In general, said solution is of light chestnut colour and viscous. To reduce as much as possible the unwanted effects of the alcohol on the components of the paper, it is useful to concentrate the resulting solution as much as possible while maintaining a
35 suitable fluidity for dosing.

The concentrated solution of carbonated magnesium

di-n-propylate is diluted to the desired concentration with a chemically inert and non-toxic diluent which allows the de-acidifying reagent to be carried inside the cellulose-type material. An HFC selected from HFC 134a and HFC 227 can be used as diluent. HFC 227 is preferable. HFC 227 has a boiling point of -17.3°C at normal pressure (101.3 kPa), a liquid density of 1.417 g/cm^3 at 20°C (399.3 kPa). The surface tensions are 9.31 mN/m (at 0°C), 6.96 mN/m (at 20°C) and 4.8 mN/m (at 40°C) and its dipole moment is $0.2\pm 0.14\text{ D}$. HFC 227 has been marketed since 1991 as a substitute for R12 and R114 in sectors where these chemicals have been used as coolants. The company SOLVAY brought this product onto the market at the beginning of 1996 as, in September 1995, the European Commission for Pharmaceutical Products in Brussels (CPMP) established that HFC 227 was suitable for pharmaceutical inhalers.

Dilution of the concentrated solution of carbonated magnesium di-n-propylate is carried out with the chosen HFC, preferably HFC 227, in pressurized containers, up to de-acidification reagent concentrations [carbonated magnesium di-n-propylate] lying between 1 % and 10 % (W/V), preferably between 1 % and 10 % (W/V) and up to concentrations of n-propanol less than 10 % (V/V). In a particular and preferred embodiment of the invention, the product of the invention contains between 3.8 and 4.5 % (W/V) of carbonated magnesium di-n-propylate, between 2 and 3 % (V/V) of n-propanol, the rest of the mixture consisting of the HFC diluent chosen from HFC 227 and HFC 134a, and the product of the invention of said formulation is especially suitable for use with spray systems.

Additionally and optionally, the product of the invention may contain an inert gas, for example, nitrogen, with a view to achieving additional pressure

and propulsion.

The resulting diluted solutions of carbonated magnesium di-n-propylate may be applied by conventional methods, for example, by direct spray onto the cellulose-type material. This method of applying the de-acidification product does not require prior desiccation, under vacuum, of the cellulose-type material, as the distribution and impregnation attained using the spray is very uniform and does not lead to the accumulation of any type of deposit or residue on the cellulose-type material to be treated. In general, it is advisable to work in a fume cupboard and use means of personal protection to avoid inhaling the spray.

The product of de-acidification product of the cellulose-type material can also be carried out by means of dissolving in HFC to a known concentration, for the bulk de-acidification of books and documents, using the equipment and method described in the Spanish patent application P9600016 filed by the same applicant. The apparatus described is formed of a treatment chamber, which serves for the drying stages under vacuum prior to impregnation and recovery of solvent by evaporation-condensation.

Therefore, the invention also provides a method for de-acidification of cellulose-type material, which comprises the application, of a quantity sufficient of the product of the invention for de-acidifying the cellulose-type material to be treated. The product of the invention may be applied to the cellulose-type material to be treated by means of a bulk de-acidification system or by spray.

The product of the invention can be obtained through a procedure that comprises:

- preparing a solution of carbonated magnesium di-n-propylate in n-propanol; and

- diluting said solution by addition of an HFC diluent selected from among HFC 134a and HFC 227.

5 In a particular embodiment, the concentration of carbonated magnesium di-n-propylate in said alcoholic solution lies between 30% and 70% (W/V).

10 Additionally and optionally, the product of the invention may contain an inert gas, for example, nitrogen.

15 The following examples serve as an illustration of the present invention and should not be considered as limiting the scope thereof.

EXAMPLE 1

Preparation of carbonated magnesium di-n-propylate

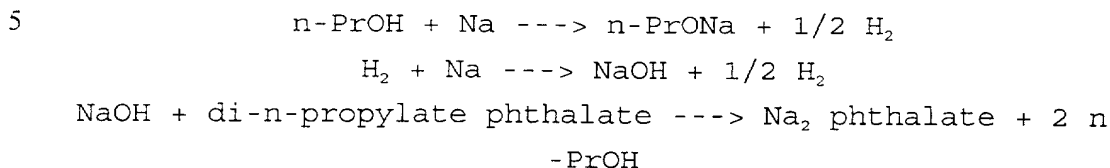
20 1.1 Dehydration of n-propanol

To dehydrate n-propanol a procedure such as the one described by A.I. Vogel, in "*Practical Organic Chemistry*", 3^a Ed. Longmans, London, 1961, pag. 168, is used, duly modified for the case of n-propanol.

25 N-propanol (Panreac quality PS) is used with a water content of less than 0.1%. The residual water is almost completely eliminated following the procedure now described.

30 In a 2-litre round-bottomed flask 1.25 l of 1-propanol are placed and 7 g of clean metallic shiny sodium added. Once the sodium has completely reacted, 27 g of di-n-ethyl phthalate are added and the mixture boiled for 2 hours. Next, n-propanol is distilled over the temperature range 97-98° C, collecting it in a dry
35 flask protected from external moisture with a tube containing anhydrous calcium chloride. This treatment allows almost complete elimination of water, given that

the sodium hydroxide formed is consumed by hydrolysis of the phthalic ester, according to the following reactions:



10

1.2 Preparation of the magnesium strip

To assure that the magnesium metal reacts well the surface layer of oxides, carbonates, etc., which has formed on coming into contact with the atmosphere, should be removed. To achieve this, about 200 g of magnesium strip are treated with 0.5 l of dilute hydrochloric acid (approximately 5 % concentration) for a short time, normally 5 minutes, shaking in an open Erlenmeyer flask so that the hydrogen is eliminated. Then it is quickly washed with distilled water several times until the acid is completely eliminated. This elimination can be verified by means of conventional methods for quantifying the presence of chloride ions. The next step is to eliminate the water with successive washes, usually 2 or 3 washes, with absolute ethanol, before drying between filter papers and storing in a topaz flask under a nitrogen atmosphere. The strip is then dried in an oven at 100° C for 15 minutes, allowed to cool, always under dry nitrogen atmosphere, and the container hermetically sealed.

1.3 Preparation of magnesium di-n-propylate

(Procedure A)

35 In a 5-litre round-bottomed flask 3.750 ml of anhydrous n-propanol are placed (3 kg, 50 mols) prepared according to the procedure described in Example 1.1.

Next, 146 g (6 mols) of magnesium metal are added in the form of a strip prepared according to the procedure described in Example 1.2 along with 3 g of iodine as catalyst. Next, the reaction mixture is heated to boiling for 6 hours, thus obtaining a suspension of magnesium di-n-propylate, a grey-coloured grey-white coloured crystalline solid in n-propanol. The resulting suspension is allowed to cool before proceeding to the next step in the reaction. A quantity of product of between 770 and 800 g is obtained, with yields of between 90 and 92 %. The concentration of magnesium di-n-propylate in said suspension lies between 34 and 35 % by weight.

15 1.4 Preparation of magnesium di-n-propylate
(Procedure B)

To 1,250 ml of anhydrous n-propanol (1 kg, 50 mols) prepared according to the procedure described in Example 1.1, 48.8 g (2 moles) of magnesium metal in the form of a strip prepared according to the procedure described in Example 1.2 are added, along with 1 g of iodine as a catalyst. Next, the reaction mixture is boiled under reflux for 16 hours, and 350-390 ml of toluene are added to the product thus obtained to eliminate part of the n-propanol and to suitably concentrate the solution to achieve product concentrations of between 45 and 50 % by weight. The toluene is used to provide an azeotropic mixture with a lower boiling point to eliminate the n-propanol. The composition of the azeotrope is 48 % propanol and 52 % toluene and the boiling point (b.p.) of the azeotrope is 92° C [b.p. of propanol: 97.2° C, and b.p. of toluene: 110.6° C].

Operating in this way, adding 390 ml of toluene leads to the elimination of 350 ml of n-propanol by the composition of the azeotrope. Thus a solid impregnated with n-propanol is obtained, with a small amount of toluene (less than 1 %), which for yields of the order

of 90 % gives a concentration of the order of 50 % of magnesium di-n-propylate.

1.5 Preparation of carbonated magnesium di-n-propylate

5 A current of carbon dioxide, completely anhydrous after passing through drying towers filled with anhydrite, is passed through a suspension of magnesium di-n-propylate in n-propanol obtained in Example 1.3 or in Example 1.4. Prior to this, said suspension of
10 magnesium di-n-propylate in n-propanol has been cooled as the reaction with carbon dioxide is exothermic and goes more slowly and with slight decomposition of the products if carried out at temperatures greater than 50° C. The solid starting product [magnesium di-n-
15 propylate] is transformed into a soluble form consisting of carbonated magnesium di-n-propylate, by raising the temperature to 45-50° C. The reaction is over after 5-6 hours, and this is manifest by a drop in the temperature of the reaction mixture to room temperature. Next, the
20 dark insoluble solid particles are allowed to sediment out until they deposit on the bottom of the flask leaving a clear and transparent solution, over a time period of 48.72 hours and the solution of carbonated magnesium di-n-propylate is collected and transferred by
25 suction under vacuum to the recipient for loading and dosing.

The amount of carbonated magnesium di-n-propylate obtained from 6 mols of magnesium is 912 g in approximately 1,700 g of n-propanol, giving rise to a
30 concentration in the carbonated product of the order of 54 % by weight.

The carbonated magnesium di-n-propylate is purified by total evaporation of n-propanol under vacuum, and;
(i) its magnesium content is analysed by complexometric
35 titration, and (ii) its bound carbon dioxide content analysed through decomposition of a sample with concentrated phosphoric acid, and collection of CO₂ in a

series of towers containing a known excess of barium hydroxide. The following results were obtained:

	Theoretical CO ₂ : 23.62%	CO ₂ found: 23.43%
5	Theoretical Mg: 13.04%	Mg found: 12.92%

The infrared spectrum of the product [carbonated magnesium di-n-propylate] shows a medium-narrow band at 536 cm⁻¹ corresponding to the Mg-O bond, and an intense band at 1,652 cm⁻¹ corresponding to the C=O bond of the carbonate.

The molar ratio Mg/CO₂ is close to 1:1, thus the formula of carbonated magnesium di-n-propylate is CH₃CH₂CH₂OMgOCOOCH₂CH₂CH₃ [MgC₇H₁₄O₄, molecular weight: 186].

The solubility of carbonated magnesium di-n-propylate in HFC 227 reaches values lying between 1 g% up to 200 g%. Thus it can be affirmed that the product is soluble at any proportion. The carbonated magnesium di-n-propylate is soluble in HFC 134a in analogue proportions.

EXAMPLE 2

Preparation of carbonated magnesium di-n-propylate

25

2.1 Preparation of magnesium di-n-propylate

(Procedure C)

In a 5-litre round-bottomed flask, 3.750 ml of anhydrous n-propanol are placed (3 kg, 50 mols), prepared according to the procedure described in Example 1.1. Next, 146 g (6 mols) of magnesium metal are added in the form of a powder with a granulometric distribution lying between 50 and 100 m, along with 7.5 g of iodine as a catalyst. Next the mixture is heated gently until hydrogen begins to be given off, and from then on, as the reaction is strongly exothermic, the reaction medium is cooled with a view to controlling the

rate of reaction at a gentle boiling for six hours until hydrogen is not longer released and the last particles of magnesium have disappeared. A suspension of magnesium di-n-propylate, a light grey-coloured solid, is obtained in n-propanol. The mixture is allowed to cool before proceeding to the next step in the reaction.

770-800 g of product are obtained, with yields between 90-92%.

2.2 Preparation of carbonated magnesium di-n-propylate

A current of carbon dioxide, completely anhydrous after passing through two drying towers full of anhydrite, is passed through the suspension of magnesium di-n-propylate in n-propanol. The reaction is exothermic, reaching temperatures of 45° C. The solid starting product [magnesium di-n-propylate] is transformed into a soluble form consisting of carbonated magnesium di-n-propylate. The reaction is over after 4-6 hours. When the temperature drops to room temperature the product is allowed to sediment out for 24-48 hours until the solid dark insoluble particles are lying at the bottom of the flask, and the solution appears clear and transparent. This solution may be collected by decantation or even better transferred by suction under vacuum, to recipients appropriate for loading, dilution or dosing. The magnesium content in this solution, measured by complexometric titration, is greater than 5 %, which represents 40 % as carbonated magnesium di-n-propylate.

The product is purified by total evaporation of n-propanol under vacuum, and the Mg content analysed by complexometric titration, and the CO₂ content is determined by decomposition of a sample with concentrated phosphoric acid and collection of CO₂ in a series of towers containing a known excess of barium hydroxide. The results obtained coincide with these mentioned in Example 1.5.

EXAMPLE 3**Efficacy assays**

In order to check the efficacy of a product for the
5 de-acidification of the cellulose-type material provided
by the invention some assays have been carried out to
determine the absorption of said product on a cellulose-
type material to assay, along with some assays to
determine the mechanical characteristics of the
10 cellulose-type material treated.

Parallel to this, by way of comparison, a
commercial paper de-acidification product [Bookkeeper]
of high efficiency has been assayed and the results
compared in each case.
15

3.1 Absorption assays

The absorption of the product for de-acidification
constitutes an indication of the alkaline reserve
created in the cellulose-type material. This reservoir
20 may play an important role in combating the degradation
of cellulose by acid hydrolysis and, therefore, is going
to contribute to the better preservation of the
cellulose material.

In the assays carried out to check the efficacy of
25 the product provided by the invention the cellulose-type
material used was paper in different stages of ageing
(without ageing and submitted to accelerated ageing).

Briefly, the assay carried out consisted of
applying the product for de-acidifying the paper to the
30 sheets of paper, either to one side or to both sides,
and then determining the alkaline reserve created in the
paper. Assays have been affected in different
conditions of ageing of the paper. The results obtained
are shown in Tables 1-3.

35

3.1.A. Without ageing

The results obtained at zero time, in other words,

without submitting the paper to an accelerated ageing process are shown in Table 1.

Table 1
Efficacy assay (Without ageing of the paper)

Sample	Initial weight (g)	Dry weight (g)	Humidity (%)	HCl	NaOH	Alkaline reserve (mol/kg)
Bookkeeper 1 side	1.017	0.948	6.78	2.565	2.39	0.184
Bookkeeper 2 sides	1.006	0.923	8.18	2.565	2.25	0.341
Invention 1 side	1.012	0.910	10.9	2.565	2.01	0.609
Invention 2 sides	1.006	0.905	10.1	2.565	1.5	1.176

Bookkeeper: Commercial product comprising micronised magnesium oxide, a humectant and a freon.

10 Invention: Solution of carbonated magnesium di-n-propylate and n-propanol in an HFC diluent selected from HFC 134a and HFC 227, at 4 %.

HCl: No. of HCl equivalents.

NaOH: No. of NaOH equivalents.

15

As can be seen, the alkaline reserve created by the product of the invention is very much better (a little greater than three times better) than that created using Bookkeeper.

20

3.1.B. Accelerated ageing (5 years)

Additional assays were carried out submitting the paper to an accelerated ageing process, consisting of keeping the paper in a chamber at a temperature of 90° C and a relative humidity of 50 %. Under these conditions, 1 hour of treatment of the paper is equivalent to an ageing of 7.5 days. Next, the product

25

is applied to the aged paper with the cellulose-type material to be de-acidified. In Table 2 the results obtained after submitting the paper to a 10-day treatment are shown. This period corresponds to an ageing of the paper of 5 years.

Table 2
Efficacy assay (Accelerated ageing of the paper: 5 years)

Sample	Initial weight (g)	Dry weight (g)	Humidity (%)	HCl	NaOH	Alkaline reserve (mol/kg)
Bookkeeper 1 side	1.015	0.962	5.57	2.565	2.40	0.171
Bookkeeper 2 sides	1.011	0.964	4.27	2.565	2.23	0.347
Invention 1 side	1.017	0.961	5.54	2.565	1.98	0.608
Invention 2 sides	1.021	0.964	5.56	2.565	1.73	0.866

Bookkeeper: Commercial product comprising micronised magnesium oxide, a humectant and a freon.

Invention: Solution of carbonated magnesium di-n-propylate and n-propanol in an HFC diluent selected from HFC 134a and HFC 227, at 4 %.

HCl: No. of HCl equivalents.

NaOH: No. of NaOH equivalents.

Just as with the previous case, the alkaline reserve created by the product of the invention is very much better (between 2.5 and 3.5 times better) than that created using Bookkeeper.

3.1.C. Accelerated ageing (10 years)

Other assays were carried out submitting the paper to the treatment for accelerated ageing as mentioned above [keeping the chamber at a temperature of 90° C and

the relative humidity at 50 %] for a period of 20 days, which represents an ageing of the paper of 10 years. Next, the product is applied to the aged paper with the cellulose-type material to be de-acidified. The results obtained are shown in Table 3.

Table 3
Efficacy assay (Accelerated ageing of the paper: 10 years)

Sample	Initial weight (g)	Dry weight (g)	Humidity (%)	HCl	NaOH	Alkaline reserve (mol/kg)
Bookkeeper 1 side	1.023	0.984	3.75	2.565	2.39	0.177
Bookkeeper 2 sides	1.028	0.977	4.95	2.565	2.32	0.250
Invention 1 side	1.022	0.972	4.84	2.565	1.83	0.756
Invention 2 sides	1.032	0.987	4.33	2.565	1.64	0.925

Bookkeeper: Commercial product comprising micronised magnesium oxide, a humectant and a freon.

Invention: Solution of carbonated magnesium di-n-propylate and n-propanol in an HFC diluent selected from HFC 134a and HFC 227, at 4 %.

HCl: No. of HCl equivalents.

NaOH: No. of NaOH equivalents.

Just as with the previous cases, but in an even more noticeable fashion, the alkaline reserve created by the product of the invention is very much better (between 3.7 and 4.2 times approximately) than that created using Bookkeeper.

3.2 Assay of mechanical properties

Through use of conventional methods (based on the corresponding ASTM, ISO and TAPPI technical guidelines)

the tearing load, the stretching, the elastic limit, the length of tear and the pH of the sheets of paper treated, in one case with the product for de-acidification of cellulose-type materials provided by this invention [Invention], and in the other case with Bookkeeper. The results obtained are shown in Table 4 where each value indicated represents the mean from 7 measurements. The pH is calculated with a mean of 5 values taken from different areas of the sheet of paper.

Table 4

Mechanical properties

[1]	[2]	[3]	[4]	[5]	[6]	[7]	[8]	[9]
Bookkeeper 1 side, 0 days	0.184	55.8 (4.2)	3.64 (0.26)	16.8 (3.5)	0.84 (0.08)	4802 (360)	54.3 (6.9)	8.3
Bookkeeper 1 side 10 days	0.171	55.4 (5.3)	2.30 (0.28)	17.4 (3.6)	0.81 (0.09)	4738 (456)	30.6 (7.8)	7.4
Bookkeeper 1 side 20 days	0.177	46.8 (5.3)	1.98 (0.28)	16.1 (8.2)	0.74 (0.32)	4070 (462)	19.9 (5.9)	8.2
Bookkeeper 2 sides 0 days	0.341	54.5 (3.2)	3.69 (0.23)	14.3 (4.1)	0.80 (0.12)	4627 (272)	53.9 (5.9)	8.5
Bookkeeper 2 sides 10 days	0.347	49.4 (3.8)	2.17 (0.25)	16.8 (2.9)	0.80 (0.06)	4268 (329)	24.0 (6.2)	7.8
Bookkeeper 2 sides 20 days	0.250	52.1 (3.0)	2.22 (0.24)	17.4 (6.2)	0.82 (0.12)	4417 (225)	26.1 (5.3)	8.0
Invention 1 side 0 day	0.609	56.3 (4.7)	3.54 (0.25)	15.9 (4.7)	0.81 (0.12)	4791 (396)	53.6 (8.8)	9.3
Invention 1 side 10 days	0.608	56.5 (8.8)	2.22 (0.45)	19.4 (5.3)	0.84 (0.11)	4736 (742)	29.3 (13)	8.0
Invention 1 side 20 days	0.756	60.3 (2.9)	2.58 (0.18)	20.1 (4.7)	0.85 (0.08)	5020 (242)	37.4 (6.0)	8.7
Invention 2 sides 0 days	1.176	65.1 (2.1)	3.42 (0.08)	18.0 (3.6)	0.81 (0.13)	5316 (171)	58.6 (3.0)	10.2
Invention 2 sides 10 days	0.866	59.6 (4.1)	2.23 (0.29)	16.2 (4.5)	0.75 (0.08)	4879 (333)	30.1 (7.3)	8.4
Invention 2 sides 20 days	0.925	54.5 (4.4)	2.38 (0.36)	21.4 (2.5)	0.89 (0.06)	4596 (373)	30.7 (8.6)	9.1

5 Key:

[1]: Sample

[2]: Alkaline reserve (mol/kg)

[3]: Tearing load (N)

[4]: Stretching (mm)

- [5]: Elastic limit (N)
[6]: Extension of elastic limit (mm)
[7]: Length of tear (m)
[8]: Traction energy absorption [T.E.A.] (J)
5 [9]: pH
Invention: Solution of carbonated magnesium di-n-propylate and n-propanol in an HFC diluent selected from HFC 134a and HFC 227, at 4 %.
Bookkeeper: Commercial product comprising micronised magnesium
10 oxide, a humectant and a freon.
The values in brackets indicate standard deviations.

The results highlight the better behaviour of the product of the invention compared to Bookkeeper.

C L A I M S

1. A product for the de-acidification of a porous material, characterised because it is presented in the form of a solution and is composed of:

- carbonated magnesium di-n-propylate,
- n-propanol, and
- and a hydroflourocarbon selected from 1,1,1,2-tetrafluoroethane (HFC 134a) and 1,1,1,2,3,3,3-heptafluoro-propane (HFC 227).

2. A product according to claim 1, characterised because the concentration of carbonated magnesium di-n-propylate in the solution formed in n-propanol and diluent lies between 1 % and 10 % (W/V).

3. A product according to claim 1, characterised because the concentration of n-propanol is less than 10 % (V/V).

4. A product according to claim 1, characterised because the concentration of carbonated magnesium di-n-propylate lies between 3.8 % and 4.5 % (W/V), the concentration of n-propanol lies between 2 % and 3 % (V/V) and the rest is made up by the diluent.

5. A procedure for obtaining a product for de-acidification of cellulose-type material according to any of claims 1 to 4, which comprises; (i) preparing a solution of carbonated magnesium di-n-propylate in n-propanol, and (ii) diluting the solution obtained in stage (i) by addition of a hydrofluorocarbon diluent selected from between HFC 134a and HFC 227.

6. A procedure according to claim 5, in which the concentration of carbonated magnesium di-n-propylate in said solution of carbonated magnesium di-n-propylate in

n-propanol lies between 30 % and 70 % (W/V).

7. A procedure according to claim 5, in which the preparation of said solution of carbonated magnesium di-n-propylate in n-propanol comprises reacting a suspension of magnesium di-n-propylate in n-propanol with dry gaseous carbon dioxide, until a solution of carbonated magnesium di-n-propylate in n-propanol is obtained; and separating the solution of carbonated magnesium di-n-propylate from in n-propanol.

8. A procedure according to claim 7, in which the preparation of said suspension of magnesium di-n-propylate in n-propanol comprises:

A) reacting magnesium metal with anhydrous n-propanol in the presence of iodine at the temperature corresponding to the boiling point of the reaction: or alternatively

B) reacting magnesium metal with anhydrous n-propanol in the presence of iodine at reflux temperature and adding toluene to form a azeotrope with n-propanol; or alternatively

C) reacting magnesium in powder form with a granulometric distribution lying between 50 and 150 m with anhydrous n-propanol in the presence of iodine and, when hydrogen begins to be released, cooling the reaction mixture to the temperature corresponding to the boiling point.

9. A method for de-acidification of cellulose-type material which comprises applying a sufficient quantity of a product for the de-acidification of cellulose material to the cellulose-type material to de-acidify according to any of the claims 1 to 4.

10. A method according to claim 9, in which said
product for de-acidification of the cellulose-type
material is applied to the cellulose-type material to be
5 treated by means of a system of de-acidification in bulk
or by spray.

PRODUCT FOR THE DE-ACIDIFICATION OF CELLULOSE MATERIAL,
AND PRODUCTION AND USE THEREOF

A B S T R A C T

5

The product is presented in the form of a solution
and is composed of carbonated magnesium di-n-propylate,
n-propanol, and a hydrofluorocarbon (HFC) diluent
selected from 1,1,1,2-tetrafluoroethane (HFC 134a) and
10 1,1,1,2,3,3,3-heptafluoropropane (HFC 227). The product
may be obtained by a procedure which comprises preparing
a solution of carbonated magnesium di-n-propylate in n-
propanol and diluting said solution by addition of the
HFC diluent. The product is suitable for de-acidifying
15 cellulose-type material and preserving objects based on
cellulose-type material.

SOLICITUD INTERNACIONAL PUBLICADA EN VIRTUD DEL TRATADO DE COOPERACION
EN MATERIA DE PATENTES (PCT)(51) Clasificación Internacional de Patentes ⁷ :

D21H 25/18

A3

(11) Número de publicación internacional:

WO 00/08250

(43) Fecha de publicación
internacional:

17 de Febrero de 2000 (17.02.00)

(21) Solicitud internacional:

PCT/ES99/00242

(22) Fecha de la presentación internacional:

29 de Julio de 1999 (29.07.99)

(30) Datos relativos a la prioridad:

P 9801733 31 de Julio de 1998 (31.07.98) ES
P 9801737 31 de Julio de 1998 (31.07.98) ES

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Asociados, S.L., Alcalá, 35, E-28014 Madrid (ES).(81) Estados designados: AE, AL, AM, AT, AU, AZ, BA, BB, BG,
BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB,
GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG,
KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK,
MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG,
SI, SK, SL, TJ, TM, TR, TT, UA, UG, US, UZ, VN, YU,
ZA, ZW, Patente ARIPO (GH, GM, KE, LS, MW, SD, SL,
SZ, UG, ZW), Patente euroasiática (AM, AZ, BY, KG, KZ,
MD, RU, TJ, TM), Patente europea (AT, BE, CH, CY, DE,
DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE),
Patente OAPI (BF, BJ, CF, CG, CI, CM, GA, GN, GW,
ML, MR, NE, SN, TD, TG).

Publicada

Con informe de búsqueda internacional.

(88) Fecha de publicación del informe de búsqueda internacional:

18 de mayo de 2000 (18.05.00)

(54) Title: PRODUCT FOR DESACIDIFICATION OF CELLULOSE MATERIAL, PRODUCTION AND UTILIZATION THEREOF

(54) Título: PRODUCTO PARA LA DESACIDIFICACION DE MATERIAL CELULOSICO, SU OBTENCION Y EMPLEO

(57) Abstract

The product is in the form of a solution and is comprised of carbonated magnesium di-n-propylate, n-propanol and a hydrofluorocarbonated diluent (HFC) selected between 1,1,1,2-tetrafluoroethane (HFC 134a) and 1,1,1,2,3,3,3-heptafluoropropane (HFC 227). The product can be obtained through a process which comprises the preparation of a solution of di-n-propylate of magnesium carbonate into n-propanol and diluting said solution by addition of the diluent HFC. The product is appropriate for desacidification of cellulose material and preservation of objects based on cellulose material.

(57) Resumen

El producto se presenta en forma de una disolución y está compuesto por di-n-propilato de magnesio carbonatado, n-propanol, y un diluyente hidrofluorocarbonado (HFC) seleccionado entre 1,1,1,2-tetrafluoroetano (HFC 134a) y 1,1,1,2,3,3,3-heptafluoro-propano (HFC 227). El producto puede obtenerse mediante un procedimiento que comprende preparar una solución de di-n-propilato de magnesio carbonatado en n-propanol y diluir dicha solución por adición del diluyente HFC. El producto es adecuado para desacidificar material celulósico y preservar objetos a base de material celulósico.

**Declaration and Power of Attorney For
Patent Application
English Language Declaration**

As a below named inventor, I hereby declare that:

My residence, post office address and citizenship are as stated below next to my name,

I believe I am the original, first and sole inventor (if only one name is listed below) or an original, first and joint inventor (if plural names are listed below) of the subject matter which is claimed and for which a patent is sought on the invention entitled PRODUCT FOR THE DE-ACIDIFICATION OF
CELLULOSE MATERIAL, AND PRODUCTION AND USE THEREOF

the specification of which

(check one)

☐ is attached hereto
☒ was filed on 1-29-01- as
Application Serial No. 091744,655-
and was amended on _____
(if applicable)

I hereby state that I have reviewed and understand the contents of the above-identified specification, including the claims, as amended by any amendment referred to above.

I acknowledge the duty to disclose information which is material to patentability as defined in Title 37, Code of Federal Regulations, § 1.56.

I hereby claim foreign priority benefits under Title 35, United States Code, §119 of any foreign application(s) for patent or inventor's certificate listed below and have also identified below any foreign application for patent or inventor's certificate having a filing date before that of the application on which priority is claimed:

Prior Foreign Application(s)			Priority Claimed	
<u>9801733</u>	<u>SPAIN</u>	<u>31 JULY 1998</u> ✓	Yes	No
(Number)	(Country)	(Day/Month/Year Filed)	<input checked="" type="checkbox"/>	<input type="checkbox"/>
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I hereby claim the benefit under Title 35, United States Code, §120 of any United States application(s) listed below, and insofar as the subject matter of each of the claims of this application is not disclosed in the prior United States application in the manner provided by the first paragraph of Title 35, United States Code, §112, I acknowledge the duty to disclose information which is material to patentability as defined in Title 37, Code of Federal Regulations, §1.56 which became available between the filing date of the prior application and the national or PCT international filing date of this application.

RECEIVED 034555 034504

(Application Serial No.) (Filing Date) (Status)
(patented, pending, abandoned)

(Application Serial No.) (Filing Date) (Status)
(patented, pending, abandoned)

I hereby declare that all statements made herein of my own knowledge are true and that all statements made on information and belief are believed to be true; and further that these statements were made with the knowledge that willful false statements and the like so made are punishable by fine or imprisonment, or both, under Section 1001 of Title 18 of the United States Code and that such willful false statements may jeopardize the validity of the application or any patent issued thereon.

POWER OF ATTORNEY: As a named inventor, I hereby appoint the following attorney(s) and/or agent(s) to prosecute this application and transact all business in the Patent and Trademark Office connected therewith.

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